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Chemistry : Physical Chemistry

Starzak, Michael E.

Energy and Entropy

Equilibrium to Stationary States

- Bridges equilibrium and time dependent processes
- Mathematical techniques are developed using simple models
- The same basic models illustrate the physical concepts in different chapters
- Temporal and spatial equations are developed using basic thermodynamics equations

The study of thermodynamics is often limited to classical thermodynamics where minimal laws and concepts lead to a wealth of equations and applications. The resultant equations best describe systems at equilibrium with no temporal or spatial parameters. The equations do, however, often provide accurate descriptions for systems close to equilibrium. Statistical thermodynamics produces the same equilibrium information starting with the microscopic properties of the atoms or molecules in the system that correlates with the results from macroscopic classical thermodynamics. Because both these disciplines develop a wealth of information from a few starting postulates, e. g., the laws of thermodynamics, they are often introduced as independent disciplines. However, the concepts and techniques developed for these disciplines are extremely useful in many other disciplines. This book is intended to provide an introduction to these disciplines while revealing the connections between them. Chemical kinetics uses the statistics and probabilities developed for statistical thermodynamics to explain the evolution of a system to equilibrium. Irreversible thermodynamics, which is developed from the equations of classical thermodynamics, centers on distance-dependent forces, and time-dependent fluxes. The force flux equations of irreversible thermodynamics lead to transport equations such as Fick's first law of diffusion and the Nernst Planck equation for electrochemical transport. The book illustrates the concepts using some simple examples.

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